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Modern Chemistry Chapter 11. Pressure. newton. barometer. millimeters of mercury. the force per unit area on a surface. the force that will increase the speed of a one-kilogram mass.... device used to measure atmospheric pressure. common unit of pressure symbolized mm Hg.

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11. An atom is said to be in an excited state when it possesses more than its minimum energy (its ground state). An atom is promoted from its ground state to an excited state by the absorption of energy, and returns to its ground state by the emission of the excess energy. 12.

Chapter 11 Modern Atomic Theory

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CHAPTER 11 REVIEW Gases Class SHORT ANSWER Answer the following questions in the space provided. c c The molar mass of a gas at STP is the density of that gas (a) multiplied by the mass of 1 mol. (b) divided by the mass of 1 mol. nRT (c) multiplied by 22.4 L. (d) divided by 22.4 L. For the expression V = (a) increasing P (b) decreasing T

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are in the atom and how they are moving—questions that we are unable to answer. The waveforms for electrons in an atom describe the variation in intensity of negative charge within the atom, with respect to the location of the nucleus. ... negative charge. Chapter 11 Modern Atomic Theory. 3 – . Chapter 11 Modern Atomic Theory. Chapter 11 ...

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Green, Michael / Modern Chemistry Textbook

Modern Chemistry Chapter 11 Flashcards | Quizlet. the rates of effusion of gases at the same temperature and pressure are inversely proportional to the square roots of their molar masses Rate of effusion of A / B = √Mb / √Ma.

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CHAPTER 11 REVIEW Gases SECTION 2 SHORT ANSWER Answer the following questions in the space provided. 1. State whether the pressure of a fixed mass of gas will increase, decrease, or stay the same in the following circumstances: increase a. temperature increases, volume stays the same decrease b. volume increases, temperature stays the same

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