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Mole-mole factors come from the coefficients of a balanced chemical equation. See Section 9.2 (Page 252) . ANSWER: Correct Chapter 9 Reading Question 2 Part A Iron reacts with oxygen in the air to produce rust (Fe_2O_3) according to the reaction shown below. $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$ How many moles of rust can be formed if 1.8 mols of iron are present? Hint 1.

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Mastering Chemistry Homework Chapter 9 Problem 9.113

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Problem 9.113 Assume that you have 1.95 g of nitroglycerin in a 605.5 mL steel container at 20.0 °C and 1.00 atm pressure. An explosion occurs, raising the temperature of the container and its contents to 425 C.

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